

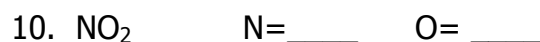
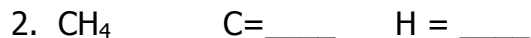
Honors Chemistry

Redox Practice

name:

hour:

For the following substances, list the oxidation number of the indicated element.



Write the oxidation number above each element in the following reaction: $\text{PbO}_2 + 4\text{HI} \rightarrow \text{I}_2 + \text{PbI}_2 + 2\text{H}_2\text{O}$

13. Which substance is oxidized? _____

14. Which substance is reduced? _____

Write the reaction for the process where

15. A neutral metal atom "M" becomes a 2+ ion

16. A 4+ metal ion becomes a 2+ metal ion

17. A neutral nonmetal atom "X" becomes a 2- ion

State whether the following represents an oxidation, a reduction, or neither for the element on both sides of the arrow

18. $\text{Na} \rightarrow \text{NaCl}$ _____

19. $\text{N}_2 \rightarrow \text{NH}_3$ _____

20. $\text{K}_2\text{Cr}_2\text{O}_7 \rightarrow \text{CrCl}_3$ _____

21. $\text{CO}_2 \rightarrow \text{CH}_4$ _____